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### Notes Of Solutions Chemistry Class

Solutions Class 12 Notes Chemistry Chapter 2. 1. A solution is a homogeneous mixture of two or 9. more chemically non-reacting substances. The components of a solution generally cannot be separated by filtration, settling or centrifuging. 2. A solution may be classified as solid, liquid or a gaseous solution. 3.

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Chemistry Notes for class 12 Chapter 2 Solutions Solution is a homogeneous mixture of two or more substances in same or different physical phases. The substances forming the solution are called components of the solution. On the basis of number of components a solution of two components is called binary solution. Solute and Solvent

### Chemistry Notes for class 12 Chapter 2 Solutions

Solutions on Solution: A solution is a homogeneous mixture of two (or more) substances, the composition of which may vary between certain limits. A solution consisting of two components is called binary solution. The component which is present in large quantity is called solvent and the component which is small in quantity is called solute.

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Revision Notes on Solution: A solution is a homogeneous mixture of two (or more) substances, the composition of which may vary between certain limits. A solution consisting of two components is called binary solution. The component which is present in large quantity is called solvent and the component which is small in quantity is called solute.

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• Saturated solutions are at dynamic equilibrium - the solute is dissolving and crystallizing at an equal rate. • The solubility of a chemical determines how much can be dissolved. Some chemicals are highly soluble (much can dissolve), some are not soluble (cannot be dissolved at all).

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Notes for Chemistry Subject for Class 12 Students are given here. Important topics of this Chapter are covered. These notes will provide you overview of chapter and important points to remember. These are very useful summary notes with neatly explained. Solution is a homogenous mixture of two or more than two compounds.

### SOLUTION HANDWRITTEN NOTES FOR CLASS 12TH CHEMISTRY

Types of solutions, expression of concentration of solutions of solids in liquids, solubility of gases in liquids, solid solutions, colligative properties - relative lowering of vapour pressure, Raoult's law, elevation of boiling point, depression of freezing point, osmotic pressure, determination of molecular masses using colligative properties, abnormal molecular mass, Van't Hoff factor.