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II. be stopped by thin paper. (ii) Give the charge on each of the radiations mentioned in 1(c) (i) above. (iii) What

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term is used to describe each of the following nuclear processes? I.

Combination of two lighter nuclei to form a heavy nucleus II. Splitting of a heavy nucleus into two or more lighter nuclei III.

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(ii) State one reason why a collision may not produce a chemical reaction. (iii) The formation of water gas is represented by the following equation. $C(s) + H_2O(g) \rightarrow CO(g) + H_2(g)$ $\Delta H = +131 \text{ kJ mol}^{-1}$. Draw an energy profile diagram for the reaction showing the . I. activated complex. II.

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(a) (i) What are acidic oxides? (ii) Give one example of each of the following oxides: I. acidic oxide; II. basic oxide; (III) amphoteric oxide; (III) neutral oxide. [5 marks] (b) (i) Define each of the following terms: (I) Heat of combustion; (II) Heat of neutralization. (ii) Write an

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equation to illustrate each of the terms in 7(b)(i) above. (iii) Given that the standard heat of combustion ...

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purification of water for town supply. (ii) Give two cations that can cause hardness in water. (iii) State two disadvantages of hardness of water. (iv) List two sources of water pollution. (b) (i) Explain briefly why water is a good solvent for sodium chloride but not for oil.

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